

SYLLABUS

Spring 2012

ENVR 403: Environmental Chemistry (3 credit hours)

Class Meets: Tuesdays & Thursdays 11:00 AM - 12:15 PM, Michael Hooker Research Center (MHRC) 03

Instructor: Dr. Jason Surratt, Assistant Professor, Atmospheric & Aerosol Chemist

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Office Hours: By appointment

Guest Lecturers:

Dr. Jason West, Assistant Professor in ESE Department, Air Pollution Scientist

Dr. Rose Cory, Assistant Professor in ESE Department, Water Chemist

Dr. Orlando Coronell, Assistant Professor in ESE Department, Civil & Environmental Engineer

Dr. Rich Kamens, Professor Emeritus in ESE Department, Atmospheric & Aerosol Chemist

Dr. Pamela Schultz, Postdoctoral Researcher at UNC & US EPA, Hydrogeochemist

Course Description and Goals: During this course you will study the chemistry of the air, water, and soil, and how anthropogenic activities affect this chemistry on planet Earth. Specifically, we will examine the sources, reactions, transport, effects, and fates of chemical species in air, water, and soil environments, and the effects of technology thereon. This course is divided into 5 major parts that reflect the most pressing issues in Environmental Chemistry today: (1) Atmospheric Chemistry and Air Pollution; (2) Climate Change and Energy; (3) Water Chemistry and Water Pollution; (4) Toxic Organic Compounds; and (5) Metals, Soils, Sediments, and Waste Disposal. All students who take this course are expected to demonstrate a mastery of all topics through successful completion of quizzes, homework sets, and exams.

Prerequisites:

This is an advanced undergraduate level course in environmental chemistry, and thus, students are required to have: (1) had at least one undergraduate course in general chemistry (2) had at least one undergraduate course in physics; (3) comfort with doing some math. It is strongly recommended that advanced undergraduates that register for this course have had at least one undergraduate course in organic chemistry. This course also serves as an introduction (or “refresher”) for first year graduate students pursuing research projects in fields related to environmental chemistry.

Grading and Course Requirements:

In-class quizzes	30%
Problem sets	20%
In-class midterm exam	20%
In-class final exam (cumulative)	30%
	100%

Final Grades Assigned:

Percentage Score	Grade
93-100	A
90-92	A-
87-89	B+
83-86	B
80-82	B-
74-79	C+
66-73	C
60-65	C-
50-59	D
< 50	F

Instructor's Philosophy of Grading Scale:

According to the UNC-CH Undergraduate Bulletin and the Registrar's website, permanent grades are defined as follows:

A Mastery of course content at the highest level of attainment that can reasonably be expected of students at a given stage of development. The A grade states clearly that the student has shown such outstanding promise in the aspect of the discipline under study that he/she may be strongly encouraged to continue.

B Strong performance demonstrating a high level of attainment for a student at a given stage of development. The B grade states that the student has shown solid promise in the aspect of the discipline under study.

C A totally acceptable performance demonstrating an adequate level of attainment for a student at a given stage of development. The C grade states that while not yet showing any unusual promise, the student may continue to study in the discipline with reasonable hope of intellectual development.

D A marginal performance in the required exercises demonstrating a minimal passing level of attainment for a student at a given stage of development. The D grade states that the student has given no evidence of prospective growth in the discipline; an accumulation of D grades should be taken to mean that the student would be well advised not to continue in the academic field.

F For whatever reasons, an unacceptable performance. The F grade indicates that the student's performance in the required exercises has revealed almost no understanding of the course content. A grade of F should warrant an adviser's questioning whether the student may suitably register for further study in the discipline before remedial work is undertaken.

For graduate students taking this course, they will be graded using the following scale:

High pass (H) – graduate students in top 10% of class will receive this grade

Pass (P) – most graduate students receive this grade if course work completed adequately

Low Pass (L)

Fail (F) – given if failure to attend lectures or labs, turn in assignments, or adequately complete quizzes/exams

In-Class Quizzes (30% of grade): Except for the first week of class, in-class quizzes will be given at the beginning of each Tuesday class period. These quizzes will be closed book and closed notes and should only last you 10 minutes (unless otherwise specified by the instructor). The purpose of the in-class quizzes is to motivate each student to review their lecture notes from the prior week as well as any required readings from the prior week so they stay up-to-speed on the course material. It is possible we may not have a quiz every Tuesday; however, you should be prepared each Tuesday regardless. I will discard your lowest quiz grade before determining your final grade for the course. If you miss a Tuesday class when a quiz is given, you will receive a grade of zero for that quiz but that grade will be the one I discard. You will also receive a grade of zero for each additional quiz missed after the first. We will likely have up to 8-10 quizzes throughout the semester, and as a result, do not worry if you miss one or do poorly on a couple of quizzes.

Problem Sets (20% of grade): Periodically (i.e., 4-6 times) during the semester, I will handout take home problem sets for you to work through. If you have difficulty with any of these assignments please feel free to schedule a time to meet with me. Problem sets are expected to be turned in on time. Late problem sets will lose 10% for each day it is late. Problem sets may be turned in late only if prior approval of the instructor is given or if there is some very exceptional circumstance that arises (e.g., severe illness, accident). The amount you learn in this course will directly relate to your ability to work problems of this level. In addition to doing the reading assignments and reviewing your lecture notes, understanding the problem sets should help prepare you for any in-class quizzes and exams.

In-Class Midterm Exam (20% of grade): The date of the in-class midterm exam is Thursday, March 1, 2012. This midterm exam will focus only on the atmospheric chemistry, air pollution, and climate change material presented through Tuesday, February 28, 2012. This exam will be closed notes and closed book. All necessary equations will be provided to you but no description of what the symbols or meaning of these equations will be provided.

In-Class Final Exam (30% of grade): The date of the in-class final exam is Tuesday, May 1, 2012 from 12:00 until 3:00 PM, in accordance with the UNC exam schedule for classes that meet TH at 11:00 AM. Unlike the midterm exam, you will have 3 hours to complete the in-class final exam. The in-class final exam will also be closed notes and closed book. The in-class final exam will cover all material discussed during the course. Students are expected to demonstrate a mastery of all major environmental chemistry concepts introduced during this course. This final exam will be held in the same classroom where we meet during the semester.

Honor Code: The University of North Carolina at Chapel Hill has had a student-administered honor system and judicial system for over 100 years. The system is the responsibility of students

and is regulated and governed by them, but faculty share the responsibility. If you have questions about your responsibility under the honor code, please bring them to your instructor or consult with the office of the Dean of Students or the Instrument of Student Judicial Governance. This document, adopted by the Chancellor, the Faculty Council, and the Student Congress, contains all policies and procedures pertaining to the student honor system. Your full participation and observance of the honor code is expected.

Students have four general responsibilities under the Honor Code:

1. Obey and support the enforcement of the Honor Code;
2. Refrain from lying, cheating, or stealing;
3. Conduct themselves so as not to impair significantly the welfare or the educational opportunities of others in the University community; and
4. Refrain from conduct that impairs or may impair the capacity of University and associated personnel to perform their duties, manage resources, protect the safety and welfare of members of the University community, and maintain the integrity of the University.

Textbooks and Other Readings:

Required Textbook:

Colin Baird and Michael Cann, Environmental Chemistry, 4th Edition, 2008.

Additional Required Readings: Research articles or photocopies of chapters from other textbooks may be handed out in class from time to time to supplement any of the material that is not adequately covered in the required textbook. Students will be required to read these in order to be better prepared for lectures, quizzes and exams.

Recommended Textbooks For Part 1 (Atmospheric Chemistry & Air Pollution) of Course for the Interested Readers:

John H. Seinfeld and Spyros Pandis, Atmospheric Chemistry and Physics: From Air pollution to Climate Change, 2nd Edition, 2006, Wiley. **[HIGHLY RECOMMENDED]**

Barbara J. Finlayson-Pitts and James N. Pitts, Jr., Chemistry of the Upper and Lower Atmosphere, 1999, Academic Press.

Recommended Textbooks or Readings for Part 2 (Climate Change & Energy) of Course for the Interested Readers:

John Houghton, Global Warming: The Complete Briefing, 4th Edition, 2009, Cambridge Univ. Press. Note: This book is used in Prof. West's class on Climate Change.

The IPCC 4th Assessment Reports (in 3 volumes), available at <http://www.ipcc.ch/>. We recommend the "Summary for Policymakers" and "Frequently Asked Questions" which communicate well to a general audience.

Recommended Textbooks For Part 3 (Water Chemistry & Water Pollution) of Course for the Interested Readers:

Brezonik, P.L.; Arnold, W.A. Water Chemistry: An Introduction to the Chemistry of Natural and Engineered Systems, Oxford University Press. 2011.

Water Quality and Treatment, 5th edition, R. Letterman, Editor, American Water Works Association, Denver, CO, 1999.

Class Schedule (subject to updates – T = Tuesdays; H = Thursdays):

Day	Day of Month	Month	Topic	Readings	Instructor
T	10	Jan	Course introduction, objectives, policies, and schedule		Surratt
H	12	Jan	PART 1: Atmospheric Chemistry and Air Pollution – Introduction to Problem	27-30, handout from Seinfeld and Pandis textbook	Surratt
T	17	Jan	Stratospheric Chemistry (Chapter 1): Physics and Chemistry of the Ozone Layer	30-46	Surratt
H	19	Jan	Stratospheric Chemistry (Chapter 1): Catalytic Processes of Ozone Destruction	47-55	Surratt
T	24	Jan	The Ozone Holes (Chapters 2, 5): The Ozone Hole and Other Sites of Ozone Depletion	59-76, 194-196	Surratt
H	26	Jan	The Ozone Holes (Chapters 2, 5): Chemicals Causing Ozone Destruction	77-87, 194-196	Surratt
T	31	Jan	Tropospheric Chemistry (Chapters 3, 5): The Photochemical Smog Process	91-131, 177-193	Surratt
H	2	Feb	Tropospheric Chemistry (Chapters 3, 5): The Photochemical Smog Process	91-131, 177-193	Surratt
T	7	Feb	Tropospheric Chemistry (Chapter 3): Particulates in Air Pollution	132-140	Surratt
H	9	Feb	Tropospheric Chemistry (Chapter 3): Particulates in	132-140	Surratt

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			Air Pollution		
T	14	Feb	Tropospheric Chemistry (Chapter 3): Particulates in Air Pollution	132-140	Surratt
H	16	Feb	Environmental and Health Effects of Polluted Air (Chapter 4): Acid Rain and Health Effects of Outdoor Air	145-171	Surratt
T	21	Feb	PART 2: Climate Change & Energy – Greenhouse Effect (Chapter 6)	203-260	West
H	23	Feb	Global Warming (Chapter 7)	261-308	West
T	28	Feb	Energy (Chapters 8-9)	311-394	West
H	1	Mar	<i>IN-CLASS MIDTERM EXAM</i>		Surratt
T	6	Mar	<i>NO CLASS – Spring Break</i>		
H	8	Mar	<i>NO CLASS – Spring Break</i>		
T	13	Mar	Part 3: Water Chemistry and Water Pollution – Chemistry of Natural Waters (Chapter 13)	555-598	Cory
H	15	Mar	Chemistry of Natural Waters (Chapter 13)	555-598	Cory
T	20	Mar	Chemistry of Natural Waters (Chapter 13)	555-598	Cory
H	22	Mar	Pollution and Purification of Water (Chapter 14)	601-651	Coronell
T	27	Mar	Pollution and Purification of Water (Chapter 14)	601-651	Coronell
H	29	Mar	Part 4: Toxic Organic Compounds – Pesticides (Chapter 10)	413-464	Kamens
T	3	Apr	Dioxins, Furans, and PCBs (Chapter 11)	469-503	Kamens
H	5	Apr	Other Toxic Organic Compounds of Environmental Concern (Chapter 12)	507-534	Kamens
T	10	Apr	Part 5: Metals, Soils, Sediments, and Waste Disposal – Toxic Heavy Metals (Chapter 15)	661-708	Schultz
H	12	Apr	Wastes, Soils, and Sediments (Chapter 16)	713-765	Schultz
T	17	Apr	Wastes, Soils, and Sediments	713-765	Schultz

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			(Chapter 16)		
H	19	Apr	Analytical Techniques used In Environmental Chemistry		Surratt
T	24	Apr	Analytical Techniques used In Environmental & Review for Final Exam		Surratt
H	1	May	<i>IN-CLASS FINAL EXAM</i>		Surratt